

Name _____

Naming Compounds Practice 1

1. Classify each of the following as an atomic element, molecular element, ionic compound, or molecular compound.

- a. helium _____
- b. PtCl_2 _____
- c. oxygen _____
- d. chlorine _____
- e. CO_2 _____
- f. cobalt _____
- g. SO_3 _____
- h. NaF _____

2. Write the formulas for the following compounds based on their name:

Name	Ions	Formula
a. Calcium chloride	_____	_____
b. Sodium phosphide	_____	_____
c. Titanium (I) sulfide	_____	_____
d. Nickel (II) perfluorate	_____	_____
e. Potassium sulfate	_____	_____
f. Zinc sulfite	_____	_____
g. Chromium (II) fluoride	_____	_____
h. Lithium iodide	_____	_____
i. Copper (II) chloride	_____	_____
j. Cobalt (II) hydroxide	_____	_____
k. Chromium (III) sulfate	_____	_____
l. Iron (III) bromide	_____	_____
m. Aluminum oxide	_____	_____
n. Magnesium hypochlorite	_____	_____

3. Write the names of the following compounds:

Name	Formula
a. CuF_2	_____
b. $\text{Mg}_3(\text{PO}_3)_2$	_____
c. Fe_2O_3	_____
d. NH_4Cl	_____
e. Ca_2P_3	_____
f. CuClO	_____
g. $\text{NaC}_2\text{H}_3\text{O}_2$	_____
h. $\text{Zn}(\text{NO}_3)_2$	_____
i. NiH_2	_____
j. MgCO_3	_____
k. CoO	_____
l. $\text{Ti}(\text{NO}_2)_2$	_____
m. KBrO_2	_____
n. BeI_2	_____
o. CuSO_4	_____
p. Cu_2SO_4	_____
q. Co_2S	_____