

Chem 20 Test 5 Review

Tro 5th Edition

You will be Given

- A periodic table
- Electronegativity chart for polarity
- Conversion factors such as 1000 mL = 1L etc

Memorize

Molarity= moles/Liters

pH= $-\log[\text{H}_3\text{O}^+]$

$C_1 \times V_1 = C_2 \times V_2$

$[\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14}$

Chapter 13 Sections 1-7

1. Define the terms solute, solvent, and solution # 1, 3, 25, 27
2. Define the terms saturated, supersaturated, and unsaturated; miscible and immiscible #6, 31, 33
3. Know the three factors increase solubility- temperature, pressure and “like dissolves like”
4. Given grams and volume, solve for molarity # 59, 61, 63
5. Given molarity and volume, solve for grams # 65, 71, 73, 75
6. Do calculations involving diluting a solution # 81, 83, 85, 87
7. Do calculations involving percent by mass # 41, 43
8. Do stoichiometry problems with solutions # 91, 93, 95

Chapter 14 Sec 1-6, 8, 9

1. Know the properties of acids and bases # 5, 7
2. Be able to do titration calculations # 47, 49
3. Correctly write neutralization reactions # 11, 39
4. Be able to do calculations with pH, pOH, H_3O^+ and OH^- # 61, 63, 65, 67, 69, 71, 73, 75, 101, 103