Chem 20 Test 5 Review

Tro 4th Edition

You will be Given

- A periodic table
- Electronegativity chart for polarity
- Conversion factors such as 1000 mL = 1L etc

Memorize

Molarity= moles/Liters $pH = -\log[H_3O^+]$

$C_1 \times V_1 = C_2 \times V_2$ $[H_3O^+][OH^-] = 1 \times 10^{-14}$

Chapter 13 Sections 1-7
1. Define the terms solute, solvent, and solution # 1, 3, 25, 27
2. Define the terms saturated, supersaturated, and unsaturated; miscible and immiscible #6, 31, 33
3. Know the three factors increase solubility- temperature, pressure and "like dissolves like"
4. Given grams and volume, solve for molarity # 59, 61, 63
5. Given molarity and volume, solve for grams # 65, 71, 73, 75
6. Do calculations involving diluting a solution # 81, 83, 85, 87
7. Do calculations involving percent by mass #41, 43
8. Do stoichiometry problems with solutions # 91, 93, 95

Chapter 14 Sec1-6, 8, 9
1. Know the properties of acids and bases #5, 7
2. Be able to do titration calculations #51, 53
3. Correctly write neutralization reactions #11, 43
4. Be able to do calculations with pH, pOH, H_3O^+ and OH^- # 65, 67, 69, 71, 73, 75, 77, 79, 107, 109